THE OCTANT RULE VI.¹ CIRCULAR DICHROISM OF (1S)-4^e-DEUTERIOADAMANTAN-2-ONE.

David A. Lightner*, Tsung C. Chang and Joseph Horwitz² Department of Chemistry, University of Nevada Reno, Nevada 89557 U.S.A.

(Received in USA 21 June 1977; received in UK for publication 7 July 1977)

Application of the Octant Rule^{1,3} to interpret $n-\pi^*$ Cotton effects (CE's) of chiral ketones with deuterium perturbers⁴ has received very little experimental investigation aside from the recent work of Djerassi *et al.*⁴ on chiral (3*R*)-deuteriocyclopentanone. The latter, despite its conformational mobility, exhibited negative circular dichroism (CD) CE's ($\Delta\varepsilon_{304} = -0.019$ at 25°C and $\Delta\varepsilon_{302} = -0.021$ at -196°C). The observed CE's are doubtless weighted averages over the contributions of all conformations of the 3-deuteriocyclopentanone and can reflect both deuterium as well as ring atom contributions.⁵ In order to establish the Octant Rule contribution due only to a deuterium perturber, we chose to prepare a stereochemically rigid, chiral molecule, (1*S*)-4^e-deuterioadamantan-2-one (1).

Deuterium was introduced stereospecifically by reaction of the ethylene ketal of the known $(1S)-4^{e}$ -bromoadmantan-2-one $(2)^{6}$ with Masamune's reagent.⁷ Hydrolysis of the deuterioketal afforded 1 after purification by thin layer chromatography and sublimation [mass spec. at 20 eV: m/e 151 (100%) 150 (16%), m.p. 256-7°, no depression upon admixture with adamantan-2-one, and IR, v: 2155 and 2170 cm⁻¹ (C-D)]. From the mass spectrum, the sample was judged to contain 86% of 1 and 14% of admantan-2-one. Since the optical purity of starting 2 was known to be 99% from earlier work,⁸ we conclude that 1 is similarly optically pure. Its epimeric homogeniety was established as > 95% equatorial deuterium by PMR analysis of the nicely separated sets of axial and equatorial hydrogens following Eu(fod)₃ addition. This technique cleanly separates the sets of four identical axial and four identical equatorial hydrogens at C-4, 8, 9 and 10 of adamantan-2-one.⁹

The CD spectrum of 1 in isopentane at 23°C is shown in the Figure, where 1 exhibits a surprisingly strong $\Delta \varepsilon_{max} = +0.08$ at 295 nm, corrected for 14% of adamantan-2-one in the sample. An octant projection diagram³ of 1 places the D-perturber in a (-) back octant¹; whereas, the observed CE is positive. Since 3^e-perturbers (except fluorine) on cyclohexanones are known to obey the Octant Rule^{1,10} and do not lie close to octant boundaries¹, and since (15)-4^e-methyladamantan-2-one derived from the precursor of 2 exhibits the expected (-) CE ($\Delta \varepsilon_{294} = -0.51$ isooctane)⁶, we conclude that deuterium, like fluorine^{3,6} makes a sign-reversed octant contribution. The reasons for this are not immediately clear. The facts that D, like F, has an atomic refractivity less than H⁴ and the C-D bond length is shorter than C-H might offer explanations.

3019



<u>FIGURE</u>. The circular dichroism spectrum of $(1S)-4^{e}$ -deuterioadamantan-2-one measured in isopentane at 23°C on a circular dichroism instrument built by Dr. J. Horwitz² and using a CAT (32 scans). A base-line is provided at or near $\Delta \varepsilon = 0$. The pertinent CD data are: $\Delta \varepsilon_{295} = +0.082$, $\Delta \varepsilon_{305} = +0.076$, $\Delta \varepsilon_{320} = +0.045$. Corrections are made to 100% optical purity and 100% d₁.

Further studies on deuterioadamantanones and 2-deuteriobicyclo[2.2.1]heptan-7-ones are in progress. This work and a theoretical treatment (T. D. Bouman) will be reported on in detail. <u>Acknowledgements</u>. We thank the National Science Foundation and the donors of the Petroleum Research Fund of the American Chemical Society for generous support of this work.

REFERENCES

- 1. For part V, see T. D. Bouman and D. A. Lightner, J. Amer. Chem. Soc., 98, 3145 (1976).
- 2. Jules Stein Eye Institute, UCLA School of Medicine, Los Angeles, California 90024.
- W. Moffitt, R. B. Woodward, A. Moscowitz, W. Klyne and C. Djerassi, J. Amer. Chem. Soc., <u>83</u>, 4013 (1961).
- 4. For leading references see J. W. Simek, D. L. Mattern and C. Djerassi, *Tetrahedron Lett.*, 3671 (1975).
- 5. T. D. Bouman, Chem. Commun., 665 (1976).
- 6. G. Snatzke and G. Eckhardt, Tetrahedron, 24, 4543 (1968).
- 7. S. Masamune, P. A. Rossy and G. S. Bates, J. Amer. Chem. Soc., <u>95</u>, 6452 (1973). We thank Professor Masamune for experimental details on the preparation of the copper(I) complex.
- 8. D. A. Lightner and T. C. Chang, J. Amer. Chem. Soc., 96, 3015 (1974).
- 9. J. B. Stothers and C. T. Tan, Chem. Commun., 738 (1974).
- 10. D. N. Kirk and W. Klyne, J.C.S. Perkin I, 1076 (1974).